



These!! I'll Give You Some Suggested Ones, But You Don't Have To Use Them. Use What Works For You. SOLUBILITY RULES (recall That Soluble Means Will Dissolve In Water) Always Soluble (will Always Completely Dissociate In Water): Alkali Metal Ions (Li +, Na +, K +, Rb +, Cs +), NH<sub>4</sub> +, NO<sub>3</sub> ... Apr 2th, 2024.

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CHEMISTRY 12 SOLUBILITY REVIEW  
QUESTIONS  
Answers To Solubility Problem Set 2 (for Questions On The Other Side):  
1.  $3.7 \times 10^{-19}$  2.  $[\text{Ca}^{2+}] = [\text{SO}_4^{2-}] = 8.4 \times 10^{-3} \text{ M}$  3.  $[\text{Ca}^{2+}] = 2.3 \times 10^{-4} \text{ M}$ ,  $[\text{F}^-] = 4.6 \times 10^{-4}$   
4.  $8.5 \times 10^{-15} \text{ G/L}$  5.  $\text{K}_{\text{TIP}} = 5.8 \times 10^{-7}$  Ppt Forms 6.  $\text{K}_{\text{TIP}} = 9.4 \times 10^{-7}$  Ppt Forms  
7. A)  $2.5 \times 10^{-5}$  B)  $3.7 \times 10^{-9}$  8.  $\text{K}_{\text{sp}}$  (at Temperature  $X$   $^{\circ}\text{C}$ ) =  $1.1 \times 10^{-8}$ ,  $X$   $^{\circ}\text{C} > 25$   $^{\circ}\text{C}$   
9.  $1.3 \times 10^{-3} \text{ M}$  (pure Water),  $1.5 \times 10^{-4} \text{ M}$  (0.10 M  $\text{Pb}(\text{NO}_3)_2$ )  
Apr 4th, 2024

AP\* Solubility Equilibrium Free Response Questions  
(a) Write A Balanced Equation For The Solubility Equilibrium. (b) Write The Expression For The Solubility Product Constant,  $\text{K}_{\text{sp}}$ , And Calculate Its Value. (c) Calculate The PH Of The Saturated Solution Of  $\text{Fe}(\text{OH})_2$  At  $25^{\circ}\text{C}$ . (d) A 50.0-milliliter Sample Of  $3.00 \times 10^{-3}$  Molar  $\text{FeSO}_4$  Solution Is Added To 50.0 Milliliters Of  $4.00 \times 10^{-6}$  M  $\text{Na}_2\text{CO}_3$  Solution. Calculate The pH Of The Solution After The Reaction Is Complete.  
Jan 4th, 2024.

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