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 33.  $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$  Actual yield = 1.45g  
 Theoretical yield =  $5.95\text{g KClO}_4 \times \frac{1\text{mol}}{122.55\text{g}} \times \frac{2\text{mol O}_2}{1\text{mol KClO}_4} \times 32.00\text{g/mol} = 3.11\text{g O}_2$   
 $\% \text{Yield} = \frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$   
 34.  $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$  Percent purity = 62.8%? g pure  $\text{FeCO}_3$  1th, 2024.  
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3 Cr(3) When The Potassium Chlorate (KClO<sub>3</sub>) Is  
Heated, It Decomposes To Form Gases KCl And  
Oxygen (O<sub>2</sub>.) 2 KClO<sub>3</sub> 2 KCl + 3 O<sub>2</sub> C. 2th, 2024.  
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INTRODUCTION 2.2 STOICHIOMETRY ...42 Chapter 2  
Environmental Chemistry Section 2.2 Stoichiometry 43

We Already Know That There Are 44 Grams Per Mole Of CO<sub>2</sub> So We Do Not Need To Recalculate That. Two Moles Of Butane (2 Mol X 58 G/mol = 116 G) Yields 8 Moles Of Carbon Dioxide (8 Mol X 44 G/mol = 352 G CO<sub>2</sub>) , So We Can Set Up The Following Proportion:  
3th, 2024.

AP\* Chemistry Stoichiometry Stoichiometry Screencasts Are Available At

Www.apchemistrynmsi.wikispaces.com 6 Exercise 7 Molar Mass And Numbers Of Molecules Isopentyl Acetate (C<sub>7</sub>H<sub>14</sub>O<sub>2</sub>), The Compound Responsible For The Scent Of Bananas, Can Be Produced Commercially. Interestingly, Bees Release About 1μg (1 × 10<sup>-6</sup> G) Of This Compound When They Sting. The Resulting Scent Attracts 3th, 2024

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