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33. $\text{KClO}_4 \rightarrow 2\text{O}_2 + \text{KCl}$ Actual yield = 1.45g
Theoretical yield = $5.95\text{g KClO}_4 \times \frac{1\text{mol}}{122.55\text{g}} \times 2 \times \frac{1\text{KClO}_4}{4} \times 32.00\text{g} = 3.11\text{g O}_2$
% Yield = $\frac{1.45\text{g}}{3.11\text{g}} \times 100\% = 46.6\%$

34. $4\text{FeCO}_3 + \text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3 + 4\text{CO}_2$ Percent purity = 62.8%? g pure FeCO_3

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We Already Know That There Are 44 Grams Per Mole Of CO₂ So We Do Not Need To Recalculate That. Two Moles Of Butane (2 Mol X 58 G/mol = 116 G) Yields 8 Moles Of Carbon Dioxide (8 Mol X 44 G/mol = 352 G CO₂) , So We Can Set Up The Following Proportion:
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Www.apchemistrynmsi.wikispaces.com 6 Exercise 7 Molar Mass And Numbers Of Molecules Isopentyl Acetate (C₇H₁₄O₂), The Compound Responsible For The Scent Of Bananas, Can Be Produced Commercially. Interestingly, Bees Release About 1μg (1 × 10⁻⁶ G) Of This Compound When They Sting. The Resulting Scent Attracts 3th, 2024

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