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(b) Divide Molar Mass By Empirical Formula Mass Molar Mass 98.96 G Empirical Formula Mass 49.48 G = 2 = (n) (c) Multiply Empirical Formula By N Obtained Above To Get The Molecular Formula Empirical Formula = CH₂Cl, N = 2. Hence Molecular Formula Is C₂H₄Cl₂. Problem 1.3 Calculate The A 4th, 2024

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Rutherford's Model. (much Greater Than/no Different From/much Less Than.) 3th, 2024

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