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Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems) Stoichiometry Worksheet #2 (mole-mass, Mass-mole Problems) 1.  $\text{N}_2 + 2\text{O}_2 \rightarrow \text{N}_2\text{O}_4$  A. If 15.0g Of  $\text{N}_2\text{O}_4$  Was Produced, How Many Moles Of  $\text{O}_2$  Were Required?  $15.0\text{g N}_2\text{O}_4 \times \frac{1 \text{ Mol N}_2\text{O}_4}{92.0\text{g N}_2\text{O}_4} = 0.163 \text{ Mol N}_2\text{O}_4$   $0.163 \text{ Mol N}_2\text{O}_4 \times \frac{2 \text{ Mol O}_2}{1 \text{ Mol N}_2\text{O}_4} = 0.326 \text{ Mol O}_2$  B. If  $4.0 \times 10^{-3}$  Moles Of Oxygen Reacted, How Many Grams Of  $\text{N}_2$  Were Needed?  $4.0 \times 10^{-3} \text{ Mol O}_2 \times \frac{1 \text{ Mol N}_2}{2 \text{ Mol O}_2} = 2.0 \times 10^{-3} \text{ Mol N}_2$   $2.0 \times 10^{-3} \text{ Mol N}_2 \times 28 \text{ g/mol} = 0.056 \text{ g N}_2$  ... Mar 10th, 2024

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