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11th, 2024 K Is The Molal Freezing Point Depression And M Molality Is ...As $1 \text{ atm} = 1.013 \times 10^5 \text{ Pa}$, The Osmotic Pressure Is $\Pi = (0.0036 \times 1.013 \times 10^5) \text{ Pa} = 365 \text{ Pa}$
Hence $C = \Pi / RT = (365 \text{ Pa}) / ((8.314 \text{ J K}^{-1} \text{ mol}^{-1}) \times (293 \text{ K})) = 0.15 \text{ mol m}^{-3}$
As $1 \text{ m}^3 = 1000 \text{ L}$, This Corresponds To $C = 1.5 \times 10^{-4} \text{ mol L}^{-1}$ May 3th, 2024.

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