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 15.0g Of N_2O_4 Was Produced, How Many Moles Of O_2 Were Required? 15.0g N_2
 O_4 1 Mol N_2O_4 2 Mol O_2 92.0g N_2O_4 1 Mol N_2O_4 = 0.326 Mol O_2 B. If
 4.0×10^{-3} Moles Of Oxygen Reacted, How Many Grams Of N_2 Were Needed?
 4.0×10^{-3} Mol O_2 1 Mol N_2 28 ... 3th, 2024Mass To Mass Stoichiometry Problems
 Worksheet Answers SO_3 As A Conversion Factor, We Determine The Mass That This
 Number Of Moles Of SO_3 Has. The First Step Resembles The Exercises We Did In
 Section 6.4 "Mole-Mole Relationships In Chemical Reactions". As Usual, We Start
 With The Quantity We Were Given: $3.59 \text{ Mol Fe}_2\text{O}_3 \times 3 \text{ Mol SO}_3 / 1 \text{ Mol Fe}_2\text{O}_3 = 10.77$
 Mol SO_3 The Mol Fe_2O_3 Units Cancel, Leaving Mol SO_3 ... 5th, 2024Stoichiometry
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