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Chapter 17: Acid-Base Equilibria And Solubility Equilibria4) 2SO₄ That Can Be

Added To 150 ML Of 0.050 M BaCl₂ Without Causing A Precipitate To Form?
Solution: First, We Have To Examine A K_{sp} Table (e.g., Table 17.4 In The Textbook).
We Can Find That The K_{sp} For BaSO₄ Is 1.1×10^{-10} This Means That If $[Ba^{2+}][SO_4^{2-}] > K_{sp}$, We Get A Pre Jan 4th, 2024

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