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The Electronic Equation That Represents The Oxidation Reaction That Occurs Is A)  $\text{HCl} + \text{KOH} \rightarrow \text{KCl} + \text{H}_2\text{O}$  B)  $4\text{HCl} + \text{MnO}_2 \rightarrow \text{MnCl}_2 + 2\text{H}_2\text{O} + \text{Cl}_2$  C)  $2\text{HCl} + \text{CaCO}_3 \rightarrow \text{CaCl}_2 + \text{H}_2\text{O} + \text{CO}_2$  D)  $2\text{HCl} + \text{FeS} \rightarrow \text{FeCl}_2 + \text{H}_2\text{S}$  21.Which Equation Represents An Oxidation-reduction Reaction? A)M Mar 16th, 2024

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A Throws A Ball With Mass  $M = 200\text{ g}$  Toward B With A Speed  $V = 21.5\text{ m/s}$  Relative To The Ice. B Catches The Ball And Throws It Back To A With The Same Speed. After A Catches The Ball, His Speed With Respect To The Ice Is A)  $4.3 \times 10^3\text{ m/s}$  D)  $9.5 \times 10^{-2}\text{ m/s}$  B)  $4.3\text{ m/s}$  E)  $0.34\text{ m/s}$  C)  $4.8 \times 10^{-2}\text{ m/s}$  Ans: D Jun 7th, 2024

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