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$N\text{O} + n\text{e}^- \rightleftharpoons \text{NRR}$ (2.1.36) The Free Energy Of The Cell Reaction Is Given By, $\Delta G = \Delta G_0 + RT \ln \frac{a_{\text{NRR}}}{a_{\text{NO}} a_{\text{H}^+}^n}$ (2.1.38) Which Since $\Delta G = -nF E_{\text{rxn}}$ And $a_{\text{H}^+} = 1$ Becomes, $E = E_0 + \frac{RT}{nF} \ln \frac{a_{\text{NRR}}}{a_{\text{NO}}}$. (2.1.40) This Expression Is Called The Nernst Equation. This Relation Can Mar 6th, 2024

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