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4. The Sum Of The Oxidation Numbers Of All The Elements In A Polyatomic Ion Equals The Charge On The Ion. Directions: In The Following Questions, Give The Oxidation Number Of The Indicated Atoms/ion. 1. N In N 2 O 3 N3+ 2. S In H 2 SO 4 S6+ 3. C 2C0 4. C In CO C2+ 2+ 5. Na In NaCl Na+ 6. H I Apr 7th, 2024

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2. For Monoatomic Ions, The Oxidation Number Is The Same As The Charge. 3. All Metals In Group 1 Form +1 Ions And Group 2 Form +2 Ions When Found In A Compound. 4. When Found In A Compound Fluorine Has A - 1 Oxidation State. 5. Hydrogen Has A +1 Oxidation State When Bonded To A Non Feb 12th. 2024

#### **RULES FOR ASSIGNING OXIDATION NUMBERS**

6. The Algebraic Sum Of The Oxidation Numbers Of Elements In A Compound Is Zero. CaF 22 H SO 4+2+2(-1)=0 2(+1) + S+4(-2)=0 Find S=+6 For Ox. No. Of S 7. The Algebraic Sum Of The Oxidation Numbers Mar 3th, 2024

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#### **Assigning Oxidation#s Practice**

Assigning Oxidation Numbers! KEY Determine The Oxidation Numbers To ALL Of The Elements In Each Of The Compounds Or Ions Below: 1) HCl 8) H 20 2 15) LiH H = +1 H = +1 (each) Li = +1 Cl = -1 O = -1 (each) H = -1 ... Microsoft Word - Assigning Oxidation#s Practice.doc Created Date: File Size: 27KB Feb 1th, 2024

#### **AP CHEMISTRY NOTES 2-1 ASSIGNING OXIDATION ...**

ASSIGNING OXIDATION NUMBERS RULES FOR ASSIGNING OXIDATION NUMBERS: 1. The Oxidation Number Of Any Free Element (including Diatomic Elements) Is Always 0. 2. The Sum Of The Oxidation Number Of All Of The Atoms In A Compound

Is Always Equal To 0. 3. The Sum Of The Oxidation Numbers Of All Of Th Apr 11th, 2024

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